

CHEM ELECTRONEGATIVITY PRACTICE

H 2.20																	He n.a.
Li 0.98	Be 1.57											B 2.04	C 2.55	N 3.04	O 3.44	F 3.98	Ne n.a.
Na 0.93	Mg 1.31											Al 1.61	Si 1.90	P 2.19	S 2.58	Cl 3.16	Ar n.a.
K 0.82	Ca 1.00	Sc 1.36	Ti 1.54	V 1.63	Cr 1.66	Mn 1.55	Fe 1.83	Co 1.88	Ni 1.91	Cu 1.90	Zn 1.65	Ga 1.81	Ge 2.01	As 2.18	Se 2.55	Br 2.96	Kr 3.00
Rb 0.82	Sr 0.95	Y 1.22	Zr 1.33	Nb 1.60	Mo 2.16	Tc 1.90	Ru 2.20	Rh 2.28	Pd 2.20	Ag 1.93	Cd 1.69	In 1.78	Sn 1.96	Sb 2.05	Te 2.10	I 2.66	Xe 2.60
Cs 0.79	Ba 0.89	La 1.10	Hf 1.30	Ta 1.50	W 2.36	Re 1.90	Os 2.20	Ir 2.20	Pt 2.28	Au 2.54	Hg 2.00	Tl 1.62	Pb 2.33	Bi 2.02	Po 2.00	At 2.20	Rn n.a.
Fr 0.70	Ra 0.89	Ac 1.10	Rf n.a.	Db n.a.	Sg n.a.	Bh n.a.	Hs n.a.	Mt n.a.	Ds n.a.	Rg n.a.	Uub n.a.	—	Uuq n.a.	—	—	—	—

Δ EN VALUES

0.0	nonpolar covalent
0.01 – 0.40	(mostly) covalent
0.40 – 1.70	polar covalent
> 1.70	(mostly) ionic

For #1-15, using the electronegativity values...

- a) calculate the bond strength
- b) classify each bond as NONPOLAR COVALENT, MODERATELY POLAR COVALENT, VERY POLAR COVALENT, or IONIC.
- c) name each compound
- d) classify each compound

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| <ol style="list-style-type: none"> 1) Ba₃P₂ 2) Al₂O₃ 3) CaI₂ 4) H₂O 5) the Na-S bond in Na₂SO₄ 6) Br₂ 7) N₂O₅ 8) RbBr | <ol style="list-style-type: none"> 9) CCl₄ 10) the Zn-C bond in Zn(CH₃COO)₂ 11) NaF 12) O₂ 13) Cu₃N 14) the K-Cl bond in KClO₃ 15) SrS |
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